

## **CHAPTER 8 & 12 ACCELERATED CHEMISTRY REVIEW SHEET**

### ***Shapes/Polarity/IMF***

*The test contains approximately 35 total questions – 22 multiple choice, 2 fill in the blank, & 11 short answer. You should study the following review sheet, your previous quiz, book problems, and practice quizzes and worksheets. Tests are weighted at 60% of your grade so be sure to prepare properly.*

**For the test you should be able to:**

- 1. Identify and determine the number of sigma and pi bonds in a molecule**
- 2. Compare the strength and length of various covalent bonds**
- 3. Identify the shapes and angles of molecules**
- 4. List the 7 – diatomic molecules**
- 5. Describe the typical bonding of carbon**
- 6. Determine the number of valence electrons an atom, ion, or molecule has**
- 7. Use a table of electronegativity's to determine the bond type**
- 8. Name covalent molecules (review)**
- 9. Draw lewis structures to determine molecular shape**
- 10. Identify the 3 exceptions to the octet rule**
- 11. Identify the type of hybrid a molecule contains**
- 12. Describe an inter and intra molecular force**
- 13. Compare the relative strength of Inter molecular forces**
- 14. Read a chart on crystalline solids and determine the properties of various solids**
- 15. Describe the octet rule**
- 16. Compare polar molecules with non-polar molecules**
- 17. Given a lewis structure predict molecular shape**
- 18. Count the number of bonded and non bonded pairs of electrons around a central atom**
- 19. Describe how the strength of intermolecular forces helps to determine the state of matter**
- 20. Describe the characteristics of a polar bond or a polar molecule**
- 21. Draw 3-dimensional shapes of molecules**